

# Gizmo Titration Answer Key

check your answers with the gizmo because dissociation is not always complete your answers may vary slightly from values in the gizmo h 2 so 4 0 ph h 2 so 4 0 hbr 0 ph hbr 0 apply for a strong base the concentration of hydroxide ions oh is estimated to be the same as the concentration of the base titration gizmo copyright all rights reserved flag for inappropriate content of 8 please do not share ruds danruddy ocsb ca titration answer key vocabulary acid analyte base dissociate equivalence point indicator litmus paper molarity neutralize ph strong acid strong base titrant titration titration curve weak acid weak base

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the titration gizmo answer key is a resource that provides the answers and explanations for the different questions and tasks in the gizmo activity it helps students check their answers and understand the concepts and techniques involved in titration

science chem is u chemistry isu name cecillia li date july 27 2021 vocabulary acid analyte base dissociate equivalence point indicator litmus paper molarity neutralize ph strong acid strong base titrant titration titration curve weak acid weak base do these before using the gizmo there are several definitions of acids

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measure a titration can be used to determine the concentration of an acid or base by measuring the amount of a solution with a known concentration called the titrant which reacts completely with a solution of unknown concentration called the analyte the point at which this occurs is called the equivalence point

cwhat is the ph of a 06 m hno<sub>3</sub> solution 2describe the equation for the reaction of nitric acid hno<sub>3</sub> and sodium hydroxide naoh is shown on the bottom right of the gizmo awhat are the reactants in this reaction bwhat are the products of this reaction 3measure a titration can be used to determine the concentration of an acid or base by

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in the titration gizmo you will use indicators to show how acids are neutralized by bases and vice versa to begin check that 100 m naoh is selected for the burette mystery hbr is selected for the flask and bromthymol blue is selected for the indicator 1 look at the flask what is the color of the bromthymol blue indicator

3 measure a titration can be used to determine the concentration of an acid or base by measuring the amount of a solution with a known concentration called the titrant which reacts completely with

a solution of unknown concentration called the analyte the point at which this occurs is called the equivalence point

check your answers with the gizmo because dissociation is not always complete your answers may vary slightly from values in the gizmo  $h_2so_4$   $ph$   $h_2so_4$   $hbr$   $ph$   $hbr$  apply for a strong base the concentration of hydroxide ions  $oh$  is roughly estimated to be the same as the concentration of the base

titration answer key vocabulary acid analyte base dissociate equivalence point indicator litmus paper molarity neutralize  $ph$  strong acid strong base titrant titration titration curve weak acid weak base quick steps to complete and design titration gizmo answer key online use get form or simply click on the template preview to open it in the editor start completing the fillable fields and carefully type in required information use the cross or check marks in the top toolbar to select your answers in the list boxes

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$d$   $160$   $m$  correct answer  $a$   $035$   $m$  the reaction of sulfuric acid  $h_2so_4$  with potassium hydroxide  $koh$  is described by the equation  $h_2so_4 + 2koh \rightarrow k_2so_4 + 2h_2o$  suppose  $50$   $ml$  of  $koh$  with unknown concentration is placed in a flask with bromthymol blue indicator a solution of  $010$   $m$   $h_2so_4$  is dripped into the  $koh$  solution

because dissociation is not always complete your answers may vary slightly from values in the gizmo  $h_2so_4$   $0$   $m$   $ph$   $h_2so_4$   $0$   $hbr$   $0$   $m$   $ph$   $hbr$   $0$  apply for a strong base the concentration of hydroxide ions  $oh$  is roughly estimated to be the same as the concentration of the base

gizmo warm up litmus is an example of an indicator a substance that changes color depending on its  $ph$  in the titration gizmo you will use indicators to show how acids are neutralized by bases and vice versa to begin check that  $100$   $m$   $naoh$  is selected for the burette mystery  $hbr$  is selected for the flask and bromthymol blue is selected

titration measure the quantity of a known solution needed to neutralize an acid or base of unknown concentration use this information to calculate the unknown concentration a variety of indicators can be used to show the  $ph$  of the solution

check your answers with the gizmo  $ca$   $oh$   $2$   $0021$   $m$   $ph$   $ca$   $oh$   $2$   $1232$   $naoh$   $0090$   $m$   $ph$   $naoh$   $1295$

titrant analyte indicator titrant volume analyte concentration  $070$   $m$   $koh$   $hbr$  vary  $30$   $ml$   $0210$   $m$   $050$   $m$   $hcl$   $ca$   $oh$   $2$  vary  $84$   $ml$   $0021$   $m$   $080$   $m$   $h_2so_4$   $naoh$  vary  $56$   $ml$   $0090$   $m$

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